

1930

The Adsorption of Chromate Ions by Colloidal Aluminum Hydroxide

Ben H. Peterson
Coe College

Kieth H. Storks
Coe College

Let us know how access to this document benefits you

Copyright ©1930 Iowa Academy of Science, Inc.

Follow this and additional works at: <https://scholarworks.uni.edu/pias>

Recommended Citation

Peterson, Ben H. and Storks, Kieth H. (1930) "The Adsorption of Chromate Ions by Colloidal Aluminum Hydroxide," *Proceedings of the Iowa Academy of Science*, 37(1), 224-224.

Available at: <https://scholarworks.uni.edu/pias/vol37/iss1/39>

This Research is brought to you for free and open access by the Iowa Academy of Science at UNI ScholarWorks. It has been accepted for inclusion in Proceedings of the Iowa Academy of Science by an authorized editor of UNI ScholarWorks. For more information, please contact scholarworks@uni.edu.

magnesium sulphate was measured. The solubility in pure water was found to be 0.003695 mols Cu $(10_3)_2$ per liter. The activity coefficients of copper iodate were calculated from this data by the equation

$$\log f = \log f_0 - \log \frac{S}{S_0}$$

the value of $\log f_0$ being obtained by extrapolation according to the method used by Le Mer. $\log \frac{S}{S_0}$ is the logarithm of the ratio of the solubility in the salt solution and in pure water. Below values of about 0.007M the observed values for $\log f$ agree with those calculated from the Bronsted and Le Mer form of the Debye- Huckel equation

$$-\log f = 1.01 \sqrt{\mu}$$

Above 0.007M the observed values are higher than the calculated.

COE COLLEGE,
CEDAR RAPIDS, IOWA.

THE ADSORPTION OF CHROMATE IONS BY COLLOIDAL ALUMINUM HYDROXIDE

BEN H. PETERSON AND KIETH H. STORKS

A stable suspension of Aluminum Hydroxide was prepared by electro-dialysis of the precipitate formed by the reaction between Aluminum Chloride and Ammonium Hydroxide. Increasing quantities of a solution of potassium chromate solution were added to equal quantities of the suspension and the quantity adsorbed determined by titrating the residual solution with standard ferrous sulphate. At the concentrations below that required for complete coagulation, the suspensoid was removed by ultrafiltration.

The results show three distinct stages of adsorption with increasing concentrations of the coagulating ion; one before complete coagulation, a second just after complete coagulation the magnitude of which seemed to bear some relation to the age of the suspension, and a third which follows the Adsorption Isotherm fairly well. It is evident that the mechanics of adsorption before and after coagulation are not identical.

COE COLLEGE,
CEDAR RAPIDS, IOWA.